



Ultra Violet Spectra Studies of Polystyrene Films in presence of Some Transition Metal Complexes with 4-amino-5-pyridyl)-4h-1,2,4-triazole-3-thiol

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ABSTRACT

Some coordination complexes of Ni(II), Cu(II), Zn(II), Cd(II) and Sn (II) metal with [4-amino-5-(pyridyl)-4H-1,2,4-triazole-3-thiol], has been prepared in alcoholic medium and used for photostabilizer for polystyrene. Polystyrene (PS) has been mixed with these complexes in chloroform solvent to form modified PS in (5%) w/v thickness, which containing concentration of complex 0.5 % by weight were produced by the casting method from chloroform solvent. The photostabilization of PS films compounds was investigated. The light of the wavelength 365 nm with intensity (6.02×10^{-9} Ein Dm⁻³ S⁻¹) is used for irradiation of PS films at room temperature. The photostabilization activity of these compounds was determined by calculating the photodecomposition rate constant (Kd) for blank and modified polystyrene.

Key words: Polystyrene, photodegradation, photostabilization, triazole.

INTRODUCTION

Triazoles are five memberd heterocyclic compounds containing three nitrogen and two carbon atoms. The substituted triazole with sulphur atom and amine group are considered to be good coordinating ligands because they involved both hard nitrogen and soft sulfur atom as thio amide group, this ligand have doner group that coordinate

with wide range of metal ions (Al-Maydama *et al.* 2008), (Narayana and Gajendragad , 1997). In this paper some complexes of substituted 1,2,4 triazol with thio and amino group has been used as photostabilizer for polystyrene. The low cost and the good performance of PS products have increased the utilization of this polymer in building and other application, it is a multipurpose polymer that is used in varied applications include rigid item.

General purpose polystyrene is clear and hard which is used in packaging, laboratory ware, and electronics (Meenakshi and Dessie 2002). One of the important uses of PS is in the manufacture of cover signal lamps of some automobiles (Safy Eldin and El-laithy, 1994), (Yousif *et al.* 2013 (A)). The photooxidative degradation processes of PS have been discussed in a number of reviews (Rabek and Ranby, 2003), (Lawrence and Weir, 2003). It is generally accepted that carbonyl formed during UV irradiation of polymers, is most probable and are responsible for the yellow coloration of the polymer (Mori *et al.* 1979). The carbonyl groups generated during the photooxidation process of polymer, extend the polymer film absorption to longer wavelengths. These groups absorb light when they irradiated with light of wavelength between (200-700 nm) and activated to the singlet and triplet excited states which enhances various successive photooxidation reactions (Cooray and Scott, 1980). The physical properties of additives and polymers play a very important role in determining the additives efficiency in photostabilization or photodegradation of polymers. For example, the compatibility that any type of additive (photostabilizer, antioxidant, thermal stabilizer. etc.) must be evenly distributed which requires that it be compatible with the polymer matrix (Gugumus, 1979), (Yousif and Haddad, 2013). In this work the photostabilization of PS was studied by using five complexes of 4-amino-5 (pyridyl)-4H-1,2,4-triazole-3-thiol.

EXPERIMENTAL

Materials

The following complexes were all prepared by the method described in this reference (Haddad *et al.*, 2013):

These complexes are:

1. Bis(4-amino-5-(pyridyl)-4H-1,2,4-triazole-3-thiol) Nickel(II), Ni(L)₂.
2. Bis(4-amino-5-(pyridyl)-4H-1,2,4-triazole-3-thiol) Copper(II), Cu(L)₂.
3. Bis(4-amino-5-(pyridyl)-4H-1,2,4-triazole-3-thiol) Zinc(II), Zn(L)₂.
4. Bis(4-amino-5-(pyridyl)-4H-1,2,4-triazole-3-thiol) Cadmium(II), Cd(L)₂.
5. Bis(4-amino-5-(pyridyl)-4H-1,2,4-triazole-3-thiol) Tin(II), Sn(L)₂.

The structure of complexes has been shown in figure (1):

Experimental techniques

Films preparation

A solution of polystyrene (5% w/v) in chloroform was used to prepare (40µm) thickness of polymer films. The films were prepared by evaporation technique at room temperature for 24 hours, to remove the possible residual chloroform solvent. After the evaporation of solvent, that's led to formation of PS films. The thickness of these films is about 40 ¼m. Modified PS can be prepared by mixing of 0.5% by weight complexes to the solution containing polystyrene (Yousif *et al.*, 2012), (Yousif *et al.*, 2013 (B))

Irradiation experiments

Accelerated Testing Technique

UV- Light was used for irradiation of PS films, giving wavelength range between (250 to 380 nm) and the maximum wavelength light intensity is at 6.2×10^{-9} Ein Dm⁻³ S⁻¹. The polymer film samples were fixed parallel to each other and the lamp of the UV. incident radiation is vertically incident on the samples. The distance between the polymer films and the source was (10 cm). (Yousif *et al.*, 2011).

Photodegradation measuring methods

Measuring the photodegradation rate of PS films by using ultraviolet-visible spectrophotometer:

The ultraviolet-visible spectrophotometer type Shimadzu UV-VIS. 160 was used to measure the changes in the Uv-visible spectrum during irradiation time for each compound at maximum absorption band (λ_{max}). The absorption spectrum was measured in the range of (200-400nm), and the (λ_{max}) at each absorption was also recorded for different irradiation times.

The infinite irradiation time was considered and the infinite absorption (A_{∞}) was assumed to be after the infinite irradiation time.

To determine the photodegradation rate constant for photostabilizer (K_d), the first order equation was used:

$$\ln(a-x) = \ln a - K_d t \quad \dots(1)$$

Where a , represent the stabilizer concentration before irradiation, x represents the change in stabilizer concentration after irradiation time (t).

If A_0 represents the absorption intensity of the polymer film containing stabilizer before irradiation, A_t represents the absorption intensity after t time of irradiation, then:

$$\begin{aligned} a &= A_0 - A_\infty \\ x &= A_0 - A_t \\ a - x &= A_0 - A_\infty - A_0 + A_t = A_t - A_\infty \dots\dots\dots (2) \end{aligned}$$

Substitution of a and $(a - x)$ from equation (2.6) in (2.5) gives:

$$\ln(A_t - A_\infty) = \ln(A_0 - A_\infty) - K_d t \quad \dots(3)$$

Thus a plot of $\ln(A_t - A_\infty)$ versus irradiation time (t) gives straight line with a slope equal (K_d) which indicates that photodecomposition of the additives is first order.

RESULTS AND DISCUSSION

Study of the efficiency of prepared Schiff's bases as photostabilizers of PS films by ultraviolet spectroscopy

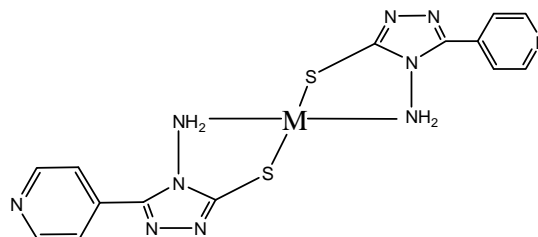
Ultraviolet radiations are known to have deleterious effects on most industrial polymers inducing chemical modification and scission of polymer chain, which ultimately lead to an undesirable loss of the mechanical and surface properties of the irradiated material. Photodegradation of PS causes discoloration (yellowing), cross-linking, and chain scission due to oxidation and effect of UV light and heat (Rabek, 1987), (Kiatkamjornwong *et al.*, 1999). UV light induces the production of radicals by oxidation. The prepared complexes were used as photostabilizers to PS films comparing with PS blank. The additives used in this study were chosen to be completely soluble in polymer solvent, (chloroform). It has been notice that the additives used in the present work are photodecomposed during the photolysis. Thus the photo decomposition rate constant (K_d) was calculated.

The K_d values were computed using the UV. spectra changes of PS films thickness 40 μ m containing 0.5% w/w from additives. The plot of irradiation time versus $\ln(A_t - A_\infty)$, gives straight line which indicate primarily the first order reaction. The slope equal to the decomposition rate constant K_d . Figure (2),(3),(4),(5),(6) and (7) shows the variation of $\ln(A_t - A_\infty)$ with irradiation time for all additives in PS films at $\lambda=365$ nm.

The values of the first order rate constant of all the modified polymers films (k_d) calculated by the same way and shown in Table (3.9).

Table 1: Photodecomposition rate constant (K_d) of PS films thickness (40mm) containing 0.5 % of additives

Polymers	K_d (S ⁻¹)
PS (blank)	4.365×10^{-7}
PS + Ni(L) ₂	7.639×10^{-8}
PS + Cu(L) ₂	2.863×10^{-8}
PS + Zn(L) ₂	3.095×10^{-8}
PS + Cd(L) ₂	1.000×10^{-7}
PS + Sn(L) ₂	2.044×10^{-7}



Where M= Ni (II), Cu (II), Zn(II), Cd(II), Sn(II)

Fig. 1: The proposed structure of Complexes

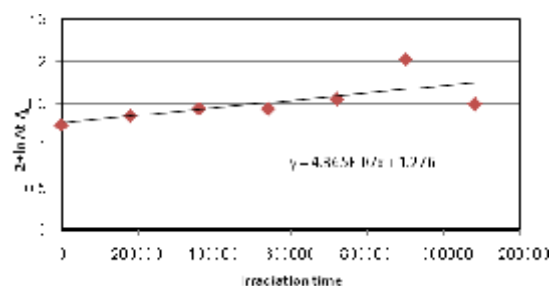


Fig. 2: Variation of natural logarithm of $\ln(A_t - A_\infty)$ with irradiation time of PS (blank) film

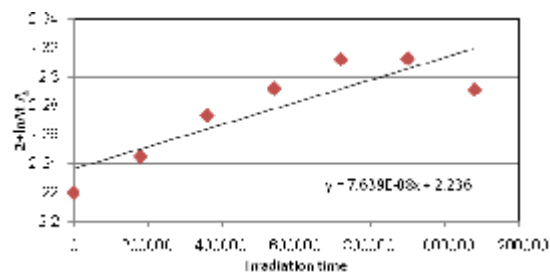


Fig. 3: Variation of natural logarithm of $\ln(A_t - A_\infty)$ with irradiation time of $\text{Ni}(\text{L})_2$ in PS film

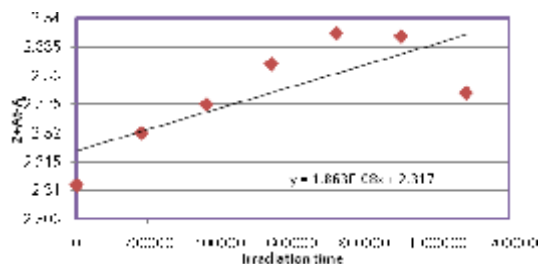


Fig. 4: Variation of natural logarithm of $\ln(A_t - A_\infty)$ with irradiation time of $\text{Cu}(\text{L})_2$ in PS film

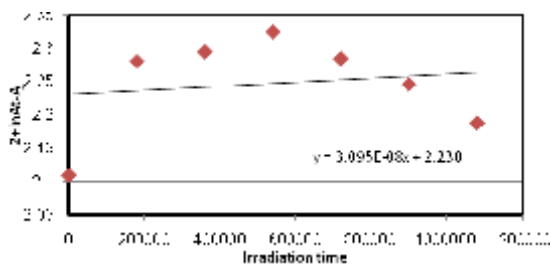


Fig. 5: Variation of natural logarithm of $\ln(A_t - A_\infty)$ with irradiation time of $\text{Zn}(\text{L})_2$ in PS film

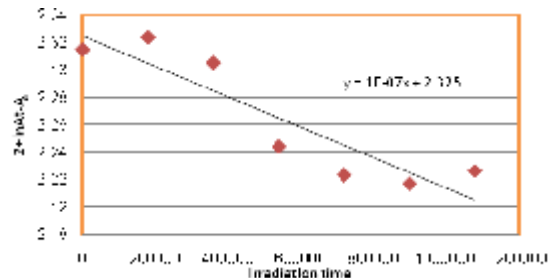


Fig. 6: Variation of natural logarithm of $\ln(A_t - A_\infty)$ with irradiation time of $\text{Cd}(\text{L})_2$ in PS film

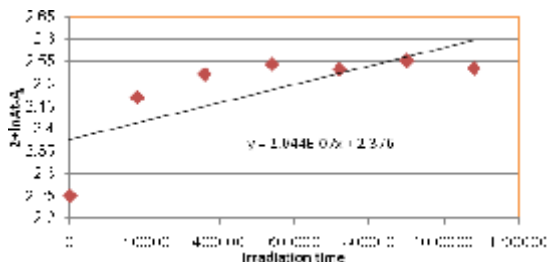
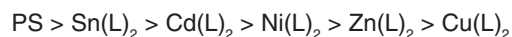


Fig. 7: Variation of natural logarithm of $\ln(A_t - A_\infty)$ with irradiation time of $\text{Sn}(\text{L})_2$ in PS film

The photostabilizers always possess low K_d values, which mean that these modified polymers are stable towards UV light. One could notice that

K_d values are sensitive to the type of additives in PS films, which decrease in the following order:



and this might point out to increase the photostability of these additives in this term.

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