



Study of the Colourimetric and Photoluminescence Properties of Polyamide fiber

AHLEM GUESMI^{1,2*} and NAOUFEL BEN HAMADI^{2,3}

¹ Textile Engineering Laboratory , Higher Institute of Technological Studies of Ksar Hellal, University of Monastir, Tunisia.

² Al Imam Mohammad Ibn Saud Islamic University (IMSIU), College of Sciences, Department of Chemistry, 11623 Riyadh, Saudi Arabia.

³ Laboratory of Synthesis Heterocyclic and Natural Substances, Faculty of Sciences of Monastir, UM (University of Monastir), Boulevard of Environment, 5000 Monastir, Tunisia.
Corresponding author E-mail : bh_naoufel@yahoo.fr

<http://dx.doi.org/10.13005/ojc/330520>

(Received: May 22, 2017; Accepted: June 20, 2017)

ABSTRACT

This research work involves the synthesis, the characterization, and the dyeing application of some reactive dyes between polyamide fibers. All the dyeing parameters have been studied, pH 4 was registered as the optimal pH value of dyeing polyamide fabrics with the synthesized dyes. The color depth as measured by the K/S value, the colorfastness to washing and light were also evaluated. Under UV exposure, intense emission of orange red was observed for dyed polyamide. Potential applications of this technology include protective clothing, textile-based sensors, and smart tagging.

Keywords: polyamide; reactive dyes; dyeing; fastness; photoluminescent.

INTRODUCTION

Many works have been interested on the dyeing of nylon fibers and many classes of dyes have been used^{1,2}. The presence of terminal amino groups in nylon fibers imparts substantivity towards anionic dyes, this made that coloration of nylon was usually achieved with acid dyes, but the major problem was related to their poor fastness properties^{3,4}. For this reason, a specific interest has been focused on the dyeing of nylon with reactive dyes, and it was known that some reactive dyes, which were developed for cellulosic fibers, can be covalently fixed at the boil, to nylon at slightly acidic

pH. Under these conditions, covalent bonds form between the dye and the amino groups of nylon, without the need for an alkaline fixation step⁵⁻⁹.

Photoluminescent polyamide fiber, that are now ubiquitous in our everyday life, have particularly attracted the attention of the scientific community these past few years due to potential important applications such as protective clothing, textile-based sensors, and smart tagging.

In the same way, we have developed the present study to elucidate the dependence between the dyeing results and the dye structure. The aim

was the synthesis of a new series of reactive dyes and to test the effect of the presence of a withdrawing electron group in the dye structure on the dyeing results. The overall dyeing parameters of nylon with synthesized dyes have been studied and optimized.

EXPERIMENTAL

Materials

Infrared spectra were recorded on a Perkin-Elmer IR-197 spectrophotometer in KBr discs. NMR spectra were obtained with a Bruker AC 300 spectrometer operating at 300 MHz for ^1H and at 75.47 MHz for ^{13}C using TMS as the internal standard. Elemental analysis was performed with a Perkin-Elmer 240 B microanalyzer. Visible absorption spectra were recorded using a Phillips model PU 8700 UV/visible spectrophotometer. Mass spectra were recorded in ESI mode. Thin layer chromatography (TLC) was performed using aluminum plates coated with silica gel 60 F254 (Merck & Co.) as stationary phase, and a mixture of isobutanol: n-propanol: ethyl acetate: water in the ratio 2:4:1:3 as mobile phase. The developed plates were visualized under both short and long wavelength ultraviolet light (254 nm, 365 nm).

Methods

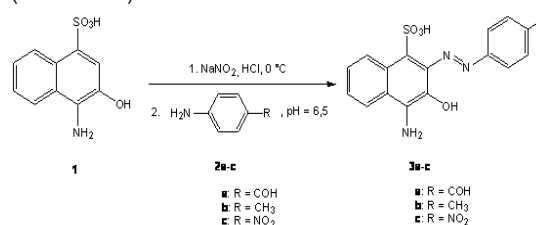
Preparation of reactive dyes

A series of dyeing experiments were directed using three novel red orange reactive dyes listed in Table 1. Dyes from 1 to 3, were prepared as described below.

Synthesis of azo dyes 3a-c

Scheme 1 represent the preparation route of azo dyes 3a-c. A solution of sodium nitrite (3.62 g, 0.058 mol), in water (20 ml), was added slowly to a mixture of aniline derivatives (0.05 mol), in water (20 ml) and hydrochloric acid (30%, 14.19 ml, 0.125 mol), at 0-5°C. After 30 min. at 0-5°C and an acidic medium (pH less than 2), excess nitrous acid was destroyed by adding sulphamic acid. The resulting solution of benzene diazonium chloride was added slowly to a solution of g-acid (1-hydroxy-7-aminonaphthalene-3-sulphonic acid) (11.9 g, 0.05 mol), at pH 6-6.5. The mixture was stirred for 30 min at room temperature and at pH 6.5. The solid product(1-hydroxy-2-arylaazo-7-aminonaphthalene-

3-sulphonicacid) was collected and purified by dissolving in distilled water (100 ml) and re-precipitating by slow addition of potassium chloride (30% w/v) to the stirred solution (Scheme 1).



Scheme 1. Synthesis of azo dyes 3a-c

Structural study of azo dyes 3a-c

Structures of all new compounds 3a-c have been elucidated by FT-IR, mass spectroscopy, elemental analyses, $^1\text{H-NMR}$ and $^{13}\text{C-NMR}$ measurements.

In IR spectra of azo dye 3b, structure was confirmed by functional group identification. Amino and alcohol group gave its peak at 3260 cm^{-1} , and appearance of the weak absorption band at 1560 cm^{-1} was due to the stretching band of azo group.

In the $^1\text{H-NMR}$ spectra of azo dye 3b, the four hydrogen atoms attached to the para-disubstituted ring gave an AA'BB' spin system.

4-Amino-2-(4-formyl-phenylazo)-3-hydroxy-naphthalene-1-sulfonic acid 3a

Solid Yield (55%); m.p. 134–135 °C. IR (KBr): $\nu/\text{cm}^{-1} = 3250$ (NH_2, OH), 1730 ($\text{C}=\text{O}$), 1550 ($\text{N}=\text{N}$). $^1\text{H NMR}$ ($\text{DMSO}-d_6$, 300 MHz): δ/ppm : 9.12 (s, 1H, $\text{C}=\text{O}$), 7.40 and 7.83 (d, 4H, $J = 8.7\text{ Hz}$, H_{arom}), 7.31-7.39 (m, 4H, H_{arom}), 10.24 (s, 2H, NH_2), 11.35 (s, 1H, OH). $^{13}\text{C NMR}$ ($\text{DMSO}-d_6$, 75.47 MHz): $\delta/\text{ppm} = 119.3, 122.7, 123.8, 125.3, 125.5, 126.6, 126.9, 127.2, 127.8, 128.7, 129.3, 130.6, 140.9, 152.7, 185.9$. Anal. Calcd for $\text{C}_{17}\text{H}_{13}\text{N}_3\text{O}_5\text{S}$: C, 54.98; H, 3.53; N, 11.31%. Found: C, 54.95; H, 3.55; N, 11.37%.

4-Amino-3-hydroxy-2-p-tolylazo-naphthalene-1-sulfonic acid 3b

Solid Yield (90%); m.p. 177–178 °C. IR (KBr): $\nu/\text{cm}^{-1} = 3270$ (NH_2, OH), 1555 ($\text{N}=\text{N}$). $^1\text{H NMR}$ ($\text{DMSO}-d_6$, 300 MHz): δ/ppm : 2.12 (s, 3H, CH_3), 6.98 and 7.39 (d, 4H, $J = 8.7\text{ Hz}$, H_{arom}), 7.42-8.04 (m, 4H, H_{arom}), 10.01 (s, 2H, NH_2), 11.58 (s, 1H, OH). $^{13}\text{C NMR}$ ($\text{DMSO}-d_6$, 75.47 MHz): $\delta/\text{ppm} = 21.51, 114.31,$

122.8, 123.6, 125.3, 126.5, 126.8 127.1, 127.6, 127.9, 129.3, 129.7, 139.8, 140.9, 149.5. Anal. Calcd for $C_{17}H_{15}N_3O_4S$: C, 57.13; H, 4.23; N, 11.76. Found: C, 57.16; H, 4.20; N, 11.81%.

4-Amino-3-hydroxy-2-(4-nitro-phenylazo)-naphthalene-1-sulfonic acid **3c**

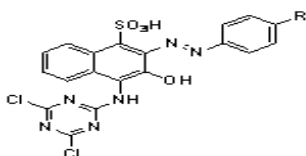
Solid Yield (75%); m.p. 146-147 °C. IR (KBr): ν/cm^{-1} = 3260 (NH₂,OH), 1560 (N=N). ¹H NMR (DMSO-*d*₆, 300 MHz): δ/ppm : 7.21-7.98 (m, 4H, H_{arom}), 8.21 and 8.37 (m, 4H, J = 7.8 Hz, H_{arom}), 10.33 (s, 2H, NH₂), 11.43 (s, 1H, OH). ¹³C NMR (DMSO-*d*₆, 75.47 MHz): δ/ppm = 119.3, 120.1, 123.2, 123.6, 124.2, 125.2, 126.3, 126.5, 127.1, 127.9, 128.7, 129.3, 140.3, 155.1. Anal. Calcd. for $C_{16}H_{12}N_4O_6S$: C, 49.48; H, 3.11; N, 14.43. Found: C, 49.39; H, 3.09; N, 14.45%.

Synthesis of reactive dyes **5a-c**

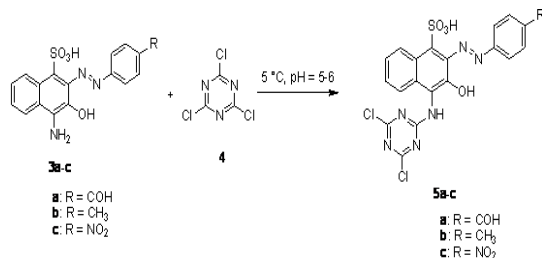
A series of dyeing experiments were conducted using three novel reactive dyes listed in Table 1. The dyes **3a-c** were prepared as described below.

Table. 1: New synthesized reactive dyes

Dye	R
5a	COH
5b	CH ₃
5c	NO ₂



2,4,6-Trichloro-s-triazine (9.4 g, 0.051 mol) in acetone (10 mL) and crushed ice (10-20 g), was added slowly to a solution of freshly purified 1-hydroxy-2-arylamino-7-aminonaphthalene-3-sulfonic acid, prepared above, in distilled water (100 mL), at below 5°C, and at pH 5-6 (Scheme 2).



Scheme. 2. Synthesis of reactive dyes **5a-c**

The mixture was stirred for 30 min. at pH 6.5 with control by TLC and the resulting red orange product, 1-hydroxy-2-arylamino-7-(2,4-dichlorotriazin-6-ylamino)-naphthalene-3-sulfonic acid, collected by filtration. For purification, the product was dissolved in distilled water (100 mL) and re-isolated by slow addition of potassium chloride (20% w/v). Due to the known instability of dichlorotriazines, the product was used direct without further characterization.

The structural of this dyes product was determined by FT-IR, mass spectroscopy, elemental analyses measurements. The band at 730 cm^{-1} was due to the stretching band of C-Cl group.

4-(4,6-Dichloro-[1,3,5]triazin-2-ylamino)-2-(4-formyl-phenylazo)-3-hydroxy naphthalene-1-sulfonic acid **5a**

Yield (95%), orange solid. m.p 300°C. Anal. Calcd. For $C_{20}H_{12}Cl_2N_6O_5S$, 46.26; H, 2.33; N, 16.18%; Found C, C, 46.23; H, 2.40; N, 16.14%. IR (KBr) ν/cm^{-1} : 730, 1550. HRMS Calcd. for $C_{20}H_{12}Cl_2N_6O_5S$ 517.9967. Found 517.9960.

4-(4,6-Dichloro-[1,3,5]triazin-2-ylamino)-3-hydroxy-2-p-tolylazo-naphthalene-1-sulfonic acid **5b**

Yield (96%), red solid. m.p 300°C. Anal. Calcd. For $C_{20}H_{14}Cl_2N_6O_4S$: C, 47.54; H, 2.79; N, 16.63%; Found C, 47.51; H, 2.83; N, 16.67%. IR (KBr) ν/cm^{-1} : 730, 1550. HRMS Calcd. for $C_{20}H_{14}Cl_2N_6O_4S$ 504.0174. Found: 504.0168.

4-(4,6-Dichloro-[1,3,5]triazin-2-ylamino)-3-hydroxy-2-(4-nitro-phenylazo)-naphthalene-1-sulfonic acid **5c**

Yield (90%), orange solid. m.p 300°C. Anal. Calcd. For $C_{19}H_{11}Cl_2N_7O_6S$, 42.55; H, 2.07; N, 18.28%; Found C, 42.59; H, 2.03; N, 18.25%. IR (KBr) ν/cm^{-1} : 730, 1550. HRMS Calcd for $C_{19}H_{11}Cl_2N_7O_6S$ 534.9869. Found 534.9858.

Dyeing

An Ahiba dyeing machine was used to dye polyamide fabric, at a liquor ratio of 20:1. Dyeing experiments were carried out at a depth of shade of 1 % o.m.f, at different pH values (2–10), for different durations (40–60 min.) and at different temperatures (80–98 °C). At the end of the dyeing, fabric was removed and rinsed in cold tap water for 5 minutes.

Colour strength

The colour yield of dyed samples was evaluated by light reflectance measurements using SF 300 spectrophotometer. Relative colour strengths (K/S values) were determined using the Kubelka–Munk equation ¹⁰.

Dye exhaustion

The extent of dye exhaustion (%E) was determined spectrophotometrically. The percentage dye bath exhaustion (%E) was calculated from Eq. (1), where A_0 and A_1 are, respectively, the absorbance of the dye bath before and after dyeing.

$$\% E = (A_0 - A_1) * 100 / A_0 \quad (1)$$

Dye fixation

In order to determine the degree of dye fixation the following procedure was used:

A sample (1 g) of dyed fabric was immersed in 150 ml of 20% aqueous pyridine solution, the temperature was raised rapidly to boiling and treatment continued under reflux until no further dye was removed from the fiber, this took approximately 3 hours. The stripped dyeing was thoroughly rinsed in tap water and allowed to dry at room temperature. The percentage of exhausted dye which fixed was determined from Eq. (2), where $(K/S)_1$ and $(K/S)_2$ represent respectively the colour strength of the dyeing fabric before and after stripping ^{11,12}.

$$\% F = (K/S)_1 / (K/S)_2 * 100 \quad (2)$$

Visible absorption spectra of dyes

Visible absorption spectra were recorded using a Philips model PU UV/visible spectrophotometer.

Fastness testing

The dyed samples were tested for fastness properties according to standard methods, the specific tests were for colour fastness to washing ISO 105-C02:1989 and colour fastness to rubbing ISO 105-X12:1987.

RESULTS AND DISCUSSION

Visible absorption spectra of dyes

The dyes are red orange and the wavelengths of maximum absorption varied from 470 to 500 nm (Tab 2). The bathochromic shift and the highest extinction coefficient of the dye 5b is due to the common effect of electron donating groups, which is CH_3 in our case ^{13,14}.

Table 2: Extinction coefficient and wavelength of maximum absorbance, of novel reactive dyes

Dye	λ_{max} (nm)	ϵ_{max} (mol ⁻¹ cm ⁻¹ L)
5a	488	20050
5b	500	22550
5c	470	19110

Effect of pH on the dyeing results

The effect of the pH value on the dyeing result have been studied. Results are collected in table 3. It was shown that the pH parameters affect enormously the dyeing results. Optimal dyeing results was registered at lowest pH value (in a pH range from 2 to 4), and then at a higher pH value, a decrease in the exhaustion level, the fixation level and also in the depth of shade have been observed.

Knowing that depending on the acidity of the medium, the polyamide fabric may be present under one of the forms present in figure 3, and that

Table 3: Effect of pH on the dyeing results. Dyeing parameters: LR 20:1; 60 min; 98°C

pH	5a				5b			5c	
	E%	E%	F%	K/S	F%	K/S	E%	F%	K/S
2	94.30	91.25	80.55	5.93	85.26	6.00	96.40	93.92	7.78
4	94.34	93.99	90.85	6.44	92.08	70.25	96.75	95.04	7.92
6	82.82	66.50	48.21	2.59	77.39	3.10	88.65	72.51	4.03
8	50.21	45.06	31.90	1.36	34.47	1.26	58.52	53.03	2.06
10	35.00	33.80	29.90	0.79	31.40	0.92	37.00	34.50	1.88

the different forms are present in an equilibrium state,¹⁵ this makes the explication of the dyeing result more trivial. Indeed, in acidic medium a nucleophilic substitution reaction occurs between the non-protonated amino end group of polyamide fibre and the chloro group of the dyes. Also, protonated amino group, which can be present in polyamide in acidic medium, can electrostatically react with the sulfonate group of the dye.

Then, increasing in pH value up to 6 promotes the formation of protonated amino end group of fiber and disadvantage the nucleophilic substitution reaction. Consecutively a decrease in a depth of shade and also in the overall dyeing parameters have been observed. At basic medium, electrostatic repulsion can be developed between the terminal carboxylate group of the fiber and the sulphonate group of the dyes, this makes the K/S value so low.

Also, comparison of the studied dyeing parameters for this series of dyes show that the highest level of the dye exhaustion and fixation was registered with the dye 5c, then 5a and then 5b. Thus, more the withdrawing electron effect of the attached group is important, more the dyeing results are better. Indeed, it was proved that the withdrawing electron group involves the charge-transfer interactions between the aryl group of the dye and carbonyl groups of the fibre. This property promotes the reaction of nucleophilic substitution and enhances consecutively the dyeing results¹⁵.

Effect of temperature on the dyeing results

The effect of the temperature of dyeing has been studied. Results are collected in table 4. It was shown that increasing the temperature of dyeing from 80 to 98°C improves the dyeing results for all the series of dyes. However, it was observed that the effect of increasing temperature was more pronounced for the dye 5b, then 5a, and then 5c and that optimal results were observed at lowest temperature for the dye 5c and 5a. Those results are reported to the effect of withdrawing electron group in those dyes.

Indeed, as previously mentioned, the withdrawing electron group involves the charge-transfer interactions between the aryl group of the dye and carbonyl groups of the fiber which facilitates the dyeability of polyamide.

Effect of dyeing time

Dyeing time was varied from 40 to 60 minutes. Results were assembled in table 5. It was observed that withdrawing electron group in dyes 5c and 5a accelerates the dye bath exhaustion when comparing with the dye 5b. Also the charge transfer interaction developed between the aryl group of the dye and carbonyl groups of the fibre, due to the withdrawing electron group, accelerates the reaction of nucleophilic substitution and enhances the fixation level from 40 minutes.

Fastness properties of dyed fabric

The rating of fastness properties (washing, and rubbing fastness) of all the dyed polyamide fabric was about 5.

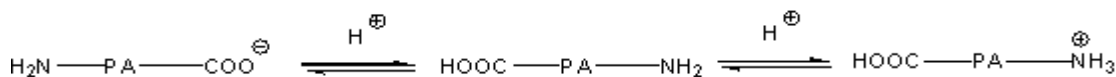


Fig. 3. Protonation's Equilibrium of nylon with varying acidity

Table. 4: Effect of temperature on the dyeing results. Dyeing parameters: LR 20:1; 60 min; pH 4

Temperature (°C)	5a			5b			5c		
	E%	F%	K/S	E%	F%	K/S	E%	F%	K/S
80	81.5	80.02	5.49	79.12	79.00	4.85	82.53	81.57	5.71
85	89.5	86.12	6.20	87.44	85.72	5.50	92.62	86.96	7.15
90	94.12	89.67	6.98	89.65	87.36	6.09	94.10	88.96	7.61
98	94.34	92.08	7.25	93.99	90.85	6.44	96.75	95.04	7.92

Luminescent on polyamide fiber

Fig. 4 shows optical image of the solid dye 5a and dyed nylon under UV light with their characteristic strong brown.

The fluorescence emission spectrum of the dyed nylon with dye 5a was determined at 488 nm excitation.

Photoluminescence spectra of the dyed nylon with dye 5a in the spectral range of 580–700 nm with excitation at 488 nm are presented in Fig. 5. A characteristic emission band of dyed nylon was observed at 625 nm with strong red orange emission color.

Table. 5: Effect of dyeing time. Dyeing parameters: LR 20:1; 98 °C; pH 4

Time (min.)	5a			5b			5c		
	E%	F%	K/S	E%	F%	K/S	E%	F%	K/S
40	92.17	89.91	6.51	85.13	81.01	5.54	94.78	91.78	7.54
50	94.06	94.50	7.05	89.85	87.56	6.03	95.8	93.56	7.80
60	94.34	92.08	7.25	93.99	90.85	6.44	96.75	95.04	7.92



Fig. 4. Intense photoluminescence of brown dye 5a on nylon was observed under UV illumination (λ 365 nm)

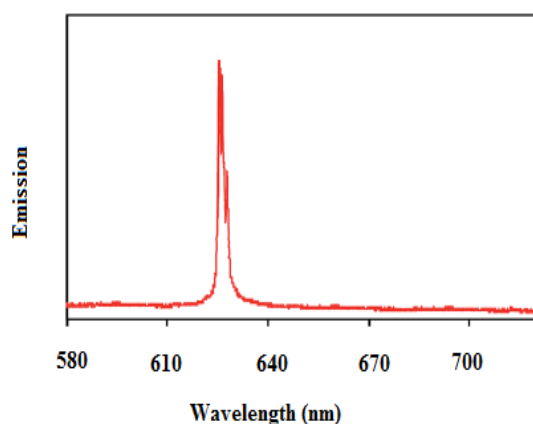


Fig. 5. Photoluminescence spectra of dye 5a grown on polyamide fiber

CONCLUSION

All the Synthesized dyes presented the highest exhaustion level on polyamide at pH 4. Withdrawing group attached to the dye structure increase the dye exhaustion and the fixation level. The study of the dyeing temperature and time has been developed, and it was shown that the withdrawing electron groups promotes the reaction of nucleophilic substitution and enhance consecutively the dyeing results at lowest time and temperature. Washing and rubbing fastness properties of all the dyed samples were good. Under UV excitation, the dyed polyamide display intense orange red emission color.

ACKNOWLEDGEMENT

This work was supported by the Textile Engineering Laboratory , Higher Institute of

Technological Studies of Ksar Hellal, University of Monastir, Tunisia.

REFERENCES

1. Burkinshaw, S.M.; Son Young, A.; Bide, M. J. *Dyes Pigments*, **2001**, *48*, 245.
2. Remington, W.R.; Gladding, E.K. *J. Am. Chem. Soc.*, **1950**, *72*, 2553.
3. Broadbent, A.D. Basic principles of textile coloration: Society of Dyers and Colourists, Bradford. **2001**.
4. Burkinshaw, S.M.; Chevli, S.N.; Marfell, D.J. *Dyes Pigments* **2000**, *45*, 235.
5. Ginns, P.; Silkstone K. In: Nunn D.M. editor. The dyeing of synthetic e polymer and acetate fibres. The Dyers Company Publications Trust **1979**.
6. Kim, S.D.; Choi, Y.J.; Lee, H.Y.; Lee, J.L. *Fibers and Polymers*, **2012**, *13*, 199.
7. Soleimani, G.A.; Taylor, J.A. *Coloration Technology*, **2011**, *127*, 227.
8. Youssef, Y.A.; Mousa, A.A.; Farouk, R.; Allam, E.E.; El-Kharadly, E.A. *Coloration Technology*, **2007**, *123*, 312.
9. Renfrew, A.H.M. Reactive dyes for textile fibres: Society of Dyers and Colorists, Bradford **1999**.
10. Judd D.B., Wysezcki, G. Colour in Business, Science and Industry, John Wiley & Sons, New York, NY, USA **1975**.
11. Lewis, D.M. Wool dyeing: Society of Dyers and Colourists, Bradford. **1992**.
12. Burkinshaw, S.M. Chemical principles of synthetic fibre dyeing: Chapman & Hall, London **1995**.
13. Uscumlic, G.S.; Petrovic, S.D. *Indian Journal of Chemistry*, **2002**, *B 41*, 206.
14. William, J.L.; Jo-Ann, H.; Banisch, M.S.; Workentin, J. *J. Chem. Soc., Chem. Commun.*, **1993**, 988.
15. Soleimani, G.A.; Taylor, J.A. *Dyes and Pigments*, **2008**, *76*, 610.